**Molecular interaction Study of Polyvinyl chloride in dimethylformamide**

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**Abstract**

In present investigation experimental values of density, viscosity and ultrasonic velocity of polyvinylchloride in dimethylformamide is measured at different concentration and temperatures at frequency of 1MHz. Using these experimental data various solute- solvent interaction parameters have been calculated to understand molecular interactions between solute and solvent. This study is helpful to understand the interaction between polymer and solvent. The change of temperature and concentration on various parameters have been studied.

**Keywords:** Ultrasonic velocity, adiabatic compressibility, Ultrasonic absorption

**I Introduction**

In recent years the ultrasonic studies of solution of polymers and in solid polymers have been the become the area of interest for research [1-5]. A literature review on volumetric studies on polymer solution reveals that measurement studies of ultrasonic velocity are helpful to understand the nature of molecular interactions in solute and solvent [6-9]. As compared to other plastics, the third most widely used plastic of synthetic polymer is Polyvinylchloride (PVC). Polyvinyl Chloride has many advantages and is generally known to be used to have advantages of least ingredient cost, wide processing adaptability, high ornamental potential and is used to produce different types of products having different range from extremely flexible to highly rapid. Using ultrasonic techniques, a revolutionary work has been done by many researchers [10-11] on polymer and polymer compatibility. As polyvinyl chloride is an important polymer for industry therefore it is chosen to study the molecular interactions of polyvinylchloride in dimethylformamide.

Takeda and Endo [12] observed the viscosity study of solution of polyvinylchloride in dilute form. B. Thapa et. al. studied the volumetric studies of polyvinylchloride with solution of cyclohexanone and 1, 4-Dioxane. During Polyvinyl chloride plastisol’s gelation and fusion with light scattering at small angle, Xiang et. al. [13] survey an effective means for the characterization of structural changes in Polyvinyl chloride particles. From the density, viscosity and viscometric measurement, the compatibility study of blend system of PVC and starch acetate in 1, 4 dioxane studied by Thakore et. al. [14]. In order to observe the effect of temperature and concentration, experiment was carried out at different concentration and temperature. Very limited literature is available on polyvinylchloride so it is interesting to study molecular interaction in polyvinylchloride. In present investigation various acoustical parameters at different range of temperature and concentration range of polyvinylchloride in dimethylformamide have been measured and the results are discussed in terms of molecular interaction of solute and solvent. Acoustical parameters are useful to understand the physio-chemical behavior of PVC and dimethyl formamide and for production, their applications in different field.

1. **Experimental Details**

## In present study Polyvinyl chloride (of molecular weight ≈ 134.5 Da) in form of liquid is used with dimethylformamide. By adding the known volume of polyvinyl chloride to fixed volume of dimethylformamide, solutions were prepared and it were stirred under reflex till a clear solution was obtained. The concentration range studied in solution is 1%, 0.8%, 0.6%, 0.5%,0.4% and 0.3% in the temperature range 30°C, 35°C, 40° C, 45° C, 50°C, 55°C, 60°C, 65°C at 1 MHz frequency. Ultrasonic velocity is measured by using variable path Ultrasonic interferometer having a reproducibility of ± 4 m/s at 30⁰C. By circulation of water by electronically operated digital constant temperature bath, the temperature of the solution has been kept constant having an accuracy of ± 0.1⁰ C) through the outer jacket of the double walled measuring cell containing experimental liquid. The density of the solution at different temperature range were measured with 10 ml specific gravity bottle and single pan micro balance. In density measurements the uncertainty was found to be 0.5 kg/m3. The viscosity of the solution was measured by the use of Ostwald’s viscometer, that was kept inside a double walled jacket, in which water was circulated from thermostat water bath. Water of desired temperature was used to fill the Inner cylinder of this double wall glass jacket so as to establish and maintain the thermal equilibrium. In uncertainty in each measurement was measured to be 0.01MPa.s. By the use of standard formulae[15-17] the different acoustical parameters are calculated.

1. **Results and discussion**

In the present investigation, the mixtures property parameters, namely, ultrasonic velocity, viscosity, density and its related parameters like adiabatic compressibility, acoustic impedance, intermolecular free length, relaxation time and ultrasonic absorption for polyvinyl chloride with dimethylformamide in different temperature range 30° C, 35°C, 40° C, 45° C, 50°C, 55°C, 60°C, 65°C and concentration range 1.0%, 0.8%, 0.6%, 0.5%,0.4% and 0.3% have been presented in Table 1 to Table 8 respectively.

The experimental values of density, viscosity and ultrasonic velocity have been presented in Table 1, 2 and 3 and their variation with temperature and concentration is presented in Figure 1 to 6 respectively. Table 1, Fig. 1 and Fig.2 indicates the variation of density with different temperature and concentration respectively. It is reported that density decreases with increase in the value of temperature and increases with increase in concentration of polyvinylchloride in the solution. Increase in the value of density is because of the reason that number of polymer chains increases in addition in the solution with increase in concentration of the PVC. As compared to solvent, polymer have large molecular weight that also contributes in the increase in density. The result reported in the present investigation are in good agreement with the results reported by earlier researchers [18, 19]. Viscosity is an important property and it depends on the molecular size, shape and intermolecular attraction. The measurement of viscosity provides useful information for study of solute- solvent interaction and for interaction of solute-solute. Table 2 and Fig.3 and Fig.4 reports the variation of viscosity with temperature and concentration respectively. It is observed that viscosity measurements decrease with temperature increase and increases with increment in concentration of polyvinylchloride in the solution. The solution becomes more and more viscous as the solute is added`, less fluidity thus led to increase in the viscosity. With the increase in the concentration, the fractional resistance between the medium layers increases and that led to increase in the viscosity. A similar behavior was made by previous authors [20]. Ultrasonic velocity decreases with increase in temperature due to weakening of intermolecular force and increases with increase in concentration of polyvinyl chloride as presented in Table 3, Fig. 5 and Fig.6 respectively. This may be due to increase in mobility of the molecules which may further increase the cohesion between molecules and thus filling all the available free spaces between it. This behaviour in the ultrasonic velocity may be because of changes in the structure in the mixture may be because of increasing intermolecular forces. Similar trend is observed by earlier authors [21]. A V Narasimham et al. have concluded the similar results for polyvinyl chloride solutions [22]. Table 4 and Fig. 7 shows that intermolecular free length increases with increase in range of temperature and decreases with increase in the concentration of Polyvinylchloride as presented in Table 4 and Fig.8. The variation of ultrasonic velocity in solution depends on intermolecular free length. According to Eyring and Kincaid model [23], ultrasonic velocity is inversely proportional to intermolecular free length, ultrasonic velocity should decrease with the increase of intermolecular free length and inverse is also true. The results reported in present study are in similar trends with proposed model. Variation of adiabatic compressibility with temperature and concentration are shown in Table 5 and Fig. 9 and 10 respectively. It is reported that adiabatic compressibility increases with increase in temperature and decreases with increase in concentration of polyvinyl chloride in solution. This decrease in value suggests the weakening of molecular interactions in the liquid mixtures. This can be explained in terms of the electrostatic effects of the polymer on the solvent molecules surrounding. The results also indicates that the medium become more compressible. The adiabatic compressibility is inversely proportional to ultrasonic velocity square value, therefore the trend in the adiabatic compressibility is reverse of the trend of ultrasonic velocity with temperature and concentration. As the velocity increases with concentration and the density does so, the compressibility must decrease with increase in concentration of PVC. Some earlier workers have also reported similar behavior of adiabatic compressibility [24]. Acoustic impedance is an important property as it is related to elastic properties of the medium. Therefore, its variation with temperature and concentration is studied. Table 6 and Fig.11 and Fig. 12 shows the variation of acoustic impedance with temperature and concentration. It is found that acoustic impedance decreases with increase in temperature and increases with increase in concentration of polyvinylchloride in solution. This is the resistance of the medium to the longitudinal wave motion. Due to the prominent interaction among the molecules, there is more resistance that results in increase in the value. This may be also due to increase in density and viscosity in solution and also because increase in elasticity of the medium. Table 7, Fig. 13 and Fig. 14 represents the variation of relaxation time with different temperature and concentration range respectively. It is observed from Table 7 and Fig.13 that relaxation time decreases with increase in the value of temperature and increases with increase in polyvinylchloride concentration in the solution (Table 7 and Fig.14). The relaxation time occurs because of the structural relaxation process and it is presumed that because of the cooperative process, the molecules get rearranged. Variation of ultrasonic absorption with temperature and concentration is presented in Table 8 and Fig. 15 and Fig.16 respectively. It is clear that ultrasonic absorption decreases with increase in temperature and increases with increase in concentration of polyvinyl chloride in solution. The increase of relaxation time and ultrasonic absorption with concentration can be explained in terms of macromolecular interchain forces motion, which are influenced by density, viscosity and ultrasonic velocity.

1. **Conclusion**

Density, viscosity and ultrasonic velocity have been measured for polyvinylchloride in dimethylformamide at different concentration and temperature. Using these values different acoustical parameters like adiabatic compressibility, acoustic impedance, intermolecular free length, relaxation time and ultrasonic absorption have been calculated. Effect of temperature and concentration have been studied on these parameters can lead to structural investigation of the medium. The results indicate that there is strong interaction between polymer and the solvent as concentration increases. This study is helpful in understanding the nature of polymer, their production and uses.

**A. Data Availability**

The data used to support finding are included within the article.

**B. Conflicts of interest**

The authors declare that they have no conflicts of interest.

1. **References**
2. S. K. Hassun, S. H. F. Al-Madfai and M. M. F. Al Jarrah, “Ultrasonic study of molecular association of poly (vinyl chloride) solution in tetrahydrofuran,” *British Polym. J*, 17(4), pp.330-333, 1985.
3. S. Bagchi, S. K. Nema and R. P. Singh, “Ultrasonic and viscometric investigation of ISRO polyol in various solvents and its compatibility with polypropylene glycol,” *European Polymer J*, 22, pp.851-857, 1986.
4. R. A. Pethrick and B. T. Poh, “Ultrasonic attenuation and adiabatic compressibility of polyethyleneoxide- water mixtures,” *The British Polymer J*, 15, pp.149-153, 1983.
5. B. Saraf and K. Samal, “Ultrasonic velocity and absorption in polystyrene solutions,” *Acustica*, 55, pp.60-63, 1984.
6. R. P. Singh, G. V. Reddy, *Acustica*, 46, pp.342, 1980.
7. O. P. Chimankar, R. Shriwas, V. A. Tabhane, “Ultrasonic absorption studies in aqueous β-Alanine and L-Glutamic Acid (Monosodium salt),” *Pegalia research library,* *Advance in Applied Research,* 1(3), pp. 78-85, 2010.
8. O. P. Chimankar, Ranjeeta S. Shriwas, Sangeeta Jajodia, and V. A. Tabhane, “Thermo- acoustic and non-linear properties of milk in NaHCO3 using volume expansion coefficient,” *Advance in Applied Research*, 2(3), pp.500 -508, 2011.
9. V. D. Bhandakkar, G. R. Bedare, V. D. Muley and B. M. Suryavanshi, “Molecular interactions of Acrylonitrile and Methacrylate in Methanol, Cyclohexane & P- dioxane,” *Advances in Applied Science Research,* 2(4) pp.338-347, 2011.
10. Kant Shashi, Kumar Sunil, Kumar Manish, Bharti Vikas, “Acoustic studies of hydrated Nickel sulphate in aqueous ascorbic acid systems,” *Der Chemica Sinica,* 3(1) , pp.166 -174, 2012.
11. R. P. Singh, Economical design of sprinkler systems, *Acoust. Soc. India,* 21, pp.159 (1993).
12. S. Rajgopalan and S J Sharma, “Study of mulberry Silk- Polyacryl amide blend using ultrasonic technique”, *J pure and Applied Ultrasonics*, 27, pp.105-109, 2005.
13. Masatami Takeda and Ryuichi Endo, “Vicosity of dilute polyvinyl chloride solution,” The *J. Physical Chem.*, 60, pp.1202-1204, 1956.
14. Guang-ming Xian, Jin-ping Qu and Bi-qing Zeng “An effective on-line Polymer characterization technique by using SALS image processing software and wavelet Analysis,” *J of Automated Methods and Management in Chemistry*, 2008, pp. 1-10, 2008.
15. I. M. Thakore, Sonal Desai,B. D. Sarwade and Surekh Devi, “Evaluation of compatibility of poly(vinyl chloride)/ starch acetate blends using simple techniques,” *J of applied Polymer Science*, 71(11)pp.1851-1861, 1999.
16. Y. Marcus, *Introduction to liquid state Chemistry,* John Wiley & Sons, London, New York, Sidney, Toronto, 1977.357 Seiten (<https://doi.org/10.1002/bbpc.197800140>)
17. Richa Saxena, S. C. Bhatt, “Molecular association studies on polyvinyl alcohol at different concentration,” *Advances in materials Science and Engineering*, 2018, pp.1-5, 2018.
18. Richa Saxena, S C Bhatt, Manish Uniyal, S C Nautiyal, “Ultrasonic and volumetric studies of aqueous solution of Polyethylene glycol,” *Journal of Mountain research*,17(2), pp.165-171, 2022.
19. B. Sundaresan and A. Srinivasa Rao, “Molecular interaction studies on poly (vinyl chloride) in chlorobenzene by acoustic method,” *Polymer Journal*, 26(10) pp. 1286-1290, 1994.
20. Gennadiy I. Egorov, Dmitriy M. Makarov, volumetric properties (water + 1,3-dimethylurea) mixture over the temperature range from 274.15 to 333.15K at the ambient pressure comparison with other methyl substituted analogues, , *J of Molecular liquids*, 323, pp.114637, 2021
21. Y. V. Patel, P. H. Parsania, “Ultrasonic velocity study of poly(R,R,4,4-cyclohexylidene diphenylene diphenyl ether-4,4-disulfonate) solutions at 30,35 and 40⁰ C,” *European Polymer Journal*, 38(10) pp.1971-1977, 2002.
22. Akanksha Saini, Aditi Prabhune, A.P. Mishra, Ranjan Dey,“Density, ultrasonic velocity, viscosity, refractive index and surface tension of aqueous choline chloride with electrolyte solutions” Journal of molecular liquids, 323, pp.114593, 2021.
23. A. V. Narasimham, S. K. J. Al-Ani, S. K. M. Al-Tayyar, “Ultrasonic properties of some polyvinyl chloride solutions,” *Indian J of Pure and applied Physics*, 37, pp.587-590,1999.
24. H. J. Erything and J. F. Kincaid, “Free volumes and free angles ratios of molecules in liquids,” *J Chem. Phys.* (USA) 6, pp.620-629, 1938.
25. D.S. Wankhede, N.N. Wankhede, M.K. Lande, B.R. Arbad, “Densities and ultrasonic velocities of some oxygen containing compounds,” *Journal of Molecular Liquids,* 138, pp. 124–129, 2008.

**Table -1: Density(x103kg/m3) of PVC at different temperature and concentration**

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **Temperature****Concentration** | **30°C** | **35°C** | **40°C** | **45°C** | **50°C** | **55°C** | **60°C** | **65°C** |
| **1.0%** | **1.053** | **1.036** | **1.031** | **0.989** | **0.979** | **0.964** | **0.948** | **0.942** |
| **0.8%** | **1.042** | **1.034** | **1.028** | **0.981** | **0.967** | **0.951** | **0.944** | **0.941** |
| **0.6%** | **0.964** | **0.937** | **0.918** | **0.867** | **0.840** | **0.831** | **0.827** | **0.825** |
| **0.5%** | **0.917** | **0.904** | **0.893** | **0.843** | **0.832** | **0.818** | **0.811** | **0.81** |
| **0.4%** | **0.891** | **0.868** | **0.855** | **0.808** | **0.796** | **0.783** | **0.777** | **0.776** |
| **0.3%** | **0.854** | **0.803** | **0.789** | **0.745** | **0.735** | **0.722** | **0.722** | **0.716** |

**Table -2: Viscosity (x10-2 MPas) of PVC at different temperature and concentration -**

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **Temperature****Concentration** | **30°C** | **35°C** | **40°C** | **45°C** | **50°C** | **55°C** | **60°C** | **65°C** |
| **1.0%** | **0.0029** | **0.0024** | **0.0022** | **0.0021** | **0.00195** | **0.0018** | **0.0017** | **0.0016** |
| **0.8%** | **0.0027** | **0.0024** | **0.0022** | **0.0021** | **0.0019** | **0.0018** | **0.0017** | **0.0016** |
| **0.6%** | **0.0027** | **0.0023** | **0.0020** | **0.0019** | **0.0017** | **0.00169** | **0.0016** | **0.0015** |
| **0.5%** | **0.0023** | **0.0021** | **0.0019** | **0.0018** | **0.0016** | **0.0015** | **0.0015** | **0.0014** |
| **0.4%** | **0.0023** | **0.0020** | **0.0019** | **0.0017** | **0.0016** | **0.0015** | **0.0015** | **0.0014** |
| **0.3%** | **0.0022** | **0.0019** | **0.0018** | **0.0017** | **0.0016** | **0.0015** | **0.0014** | **0.0013** |

**Table -3: Ultrasonic velocity (m/s) of PVC at different temperature and concentration-**

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **Temperature****Concentration** | **30°C** | **35°C** | **40°C** | **45°C** | **50°C** | **55°C** | **60°C** | **65°C** |
| **1.0%** | **1264.1** | **1258.3** | **1210.8** | **1197.5** | **1189.8** | **1160.8** | **1153.9** | **1143.6** |
| **0.8%** | **1242.8** | **1238.6** | **1207.3** | **1190.6** | **1187.6** | **1154** | **1143** | **1138.5** |
| **0.6%** | **1240.6** | **1210.4** | **1204.6** | **1188.3** | **1177.3** | **1150.6** | **1142.6** | **1136** |
| **0.5%** | **1224.2** | **1206** | **1181.6** | **1178** | **1150.7** | **1148.4** | **1142** | **1132.2** |
| **0.4%** | **1187** | **1181.8** | **1172.5** | **1165.2** | **1150.4** | **1146.7** | **1142** | **1131.6** |
| **0.3%** | **1177.6** | **1168.1** | **1156.8** | **1152** | **1146** | **1136.6** | **1131.2** | **1130.6** |

**Table – 4: Intermolecular Free Length (x10-13m) of PVC at different temperature and concentration-**

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **Temperature****Concentration** | **30°C** | **35°C** | **40°C** | **45°C** | **50°C** | **55°C** | **60°C** | **65°C** |
| **1.0%** | **3.3** | **3.341** | **3.481** | **3.596** | **3.637** | **3.756** | **3.809** | **3.855** |
| **0.8%** | **3.374** | **3.396** | **3.496** | **3.629** | **3.665** | **3.803** | **3.855** | **3.875** |
| **0.6%** | **3.514** | **3.653** | **3.708** | **3.868** | **3.967** | **4.056** | **4.123** | **4.24** |
| **0.5%** | **3.65** | **3.795** | **3.834** | **3.957** | **4.087** | **4.122** | **4.148** | **4.2** |
| **0.4%** | **3.819** | **3.887** | **3.948** | **4.087** | **4.17** | **4.218** | **4.252** | **4.266** |
| **0.3%** | **3.933** | **4.164** | **4.164** | **4.305** | **4.344** | **4.431** | **4.452** | **4.475** |

**Table – 5: Adiabatic compressibility(x10-10kg-1m4s2) of PVC at different temperature and**

**concentration-**

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **Temperature****Concentration** | **30°C** | **35°C** | **40°C** | **45°C** | **50°C** | **55°C** | **60°C** | **65°C** |
| **1.0%** | **5.95** | **6.09** | **6.62** | **7.06** | **7.22** | **7.7** | **7.92** | **8.11** |
| **0.8%** | **6.21** | **6.30** | **6.67** | **7.19** | **7.33** | **7.89** | **8.11** | **8.20** |
| **0.6%** | **6.74** | **7.28** | **7.51** | **8.17** | **8.59** | **8.98** | **9.28** | **9.58** |
| **0.5%** | **7.27** | **7.86** | **8.02** | **8.55** | **9.08** | **9.28** | **9.39** | **9.63** |
| **0.4%** | **7.96** | **8.25** | **8.51** | **9.12** | **9.49** | **9.71** | **9.87** | **9.94** |
| **0.3%** | **8.44** | **9.13** | **9.47** | **9.51** | **9.53** | **9.77** | **9.88** | **9.96** |

**Table -6: Acoustic impedance (x103kgm2s-1) of PVC at different temperature and concentration-**

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **Temperature****Concentration** | **30°C** | **35°C** | **40°C** | **45°C** | **50°C** | **55°C** | **60°C** | **65°C** |
| **1.0%** | **1330.6** | **1304.1** | **1248.2** | **1183.1** | **1164.2** | **1118.9** | **1094.0** | **1077.6** |
| **0.8%** | **1294.9** | **1282.2** | **1241.6** | **1168.2** | **1148.5** | **1097.7** | **1078.7** | **1071.7** |
| **0.6%** | **1195.8** | **1134.4** | **1105.9** | **1030.5** | **988.6** | **962.0** | **944.1** | **920.1** |
| **0.5%** | **1123.1** | **1072.4** | **1054.9** | **993.3** | **957.2** | **938.8** | **928.8** | **917.2** |
| **0.4%** | **1057.9** | **1026.0** | **1002.3** | **941.4** | **915.6** | **897.9** | **887.2** | **883.6** |
| **0.3%** | **1005.7** | **937.5** | **913.2** | **903.2** | **858.0** | **821.0** | **817.0** | **809.2** |

**Table – 7: Relaxation time (x10-12s) of PVC at different temperature and concentration-**

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **Temperature****Concentration** | **30°C** | **35°C** | **40°C** | **45°C** | **50°C** | **55°C** | **60°C** | **65°C** |
| **1.0%** | **2.720** | **2.520** | **2.460** | **2.455** | **1.970** | **1.920** | **1.838** | **1.764** |
| **0.8%** | **2.680** | **2.498** | **2.420** | **2.394** | **1.950** | **1.876** | **1.860** | **1.700** |
| **0.6%** | **2.640** | **2.420** | **2.391** | **2.351** | **1.868** | **1.830** | **1.82** | **1.976** |
| **0.5%** | **2.610** | **2.380** | **2.330** | **2.251** | **1.780** | **1.942** | **1.726** | **1.605** |
| **0.4%** | **2.560** | **2.332** | **2.290** | **2.227** | **1.750** | **1.730** | **1.690** | **1.62** |
| **0.3%** | **2.454** | **2.310** | **2.234** | **2.180** | **1.72** | **1.690** | **1.620** | **1.600** |

**Table -8: Ultrasonic absorption (x10-15s2m-1) of PVC at different temperature and concentration-**

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **Temperature****Concentration** | **30°C** | **35°C** | **40°C** | **45°C** | **50°C** | **55°C** | **60°C** | **65°C** |
| **1.0%** | **3.56** | **3.52** | **3.47** | **3.44** | **3.40** | **3.34** | **3.31** | **3.20** |
| **0.8%** | **3.55** | **3.50** | **3.45** | **3.41** | **3.38** | **3.28** | **3.22** | **3.16** |
| **0.6%** | **3.51** | **3.47** | **3.42** | **3.38** | **3.35** | **3.25** | **3.20** | **3.05** |
| **0.5%** | **3.48** | **3.45** | **3.38** | **3.35** | **3.33** | **3.21** | **3.18** | **3.03** |
| **0.4%** | **3.45** | **3.41** | **3.35** | **3.32** | **3.29** | **3.18** | **3.15** | **3.01** |
| **0.3%** | **3.42** | **3.39** | **3.32** | **3.30** | **3.25** | **3.21** | **3.14** | **3.01** |

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**Fig.-1: Variation of density with temperature at different concentration of PVC**

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**Fig.-2: Variation of density with concentration at different temperature of PVC**

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 **Fig.-3: Variation of Viscosity with temperature at different concentration of PVC**

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**Fig.-4: Variation of Viscosity with concentration at different temperature of PVC **

**Fig.-5: Variation of ultrasonic velocity with temperature at different concentration of PVC**

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**Fig.-6: Variation of ultrasonic velocity with concentration at different temperature of PVC**

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**Fig. -7: Variation of intermolecular free length with temperature at different concentration of PVC**

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**Fig.- 8: Variation of intermolecular free length with concentration at different temperature of PVC**

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**Fig. -9: Variation of adiabatic compressibility with temperature at different concentration of PVC**

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**Fig.-10: Variation of adiabatic compressibility with concentration at different temperature of PVC**

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**Fig. -11: Variation of acoustic impedance with temperature at different concentration of PVC**

****

**Fig. -12: Variation of acoustic impedance with concentration at different temperature of PVC**

**1.5**

**1.7**

**1.9**

**2.1**

**2.3**

**2.5**

**2.7**

**2.9**

**30**

**35**

**40**

**45**

**50**

**55**

**60**

**65**

**Temperature (**

**0**

**C)**

**Relaxation time(x10-12s)**

1

0.8

0.6

0.5

0.4

0.3

**Fig.-13: Variation of relaxation time with temperature at different concentration of PVC**

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**Fig.-14: Variation of relaxation time with concentration at different temperature of PVC**

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**Fig.-15: Variation of ultrasonic absorption with temperature at different concentration of PVC**



**Fig.-16: Variation of ultrasonic absorption with concentration at different temperature of PVC**